Percent Yield Practice (#7)
adapted from *ChemFiesta.com*

1)   a) Write the balanced equation:
          lithium hydroxide + potassium chloride → _____ + _____
      b) The reaction began with 20.0 grams of lithium hydroxide. What is the
          theoretical yield of lithium chloride?
      c) The reaction produced 6.00 grams of lithium chloride. What is the percent
          yield?

2)   a) Write the balanced equation:
          beryllium + hydrochloric acid → _____ + _____
      b) A student's theoretical yield of beryllium chloride was 10.7 grams. If the
          actual yield was 4.50 grams, what was the percent yield?

3)   a) Write the balanced equation:
          sodium chloride + calcium oxide → _____ + _____
      b) What is the theoretical yield of sodium oxide if a chemist begins with 20.0
          grams of calcium oxide?

4)   a) Write the balanced equation:
          iron(II) bromide + potassium chloride → _____ + _____
      b) What is the theoretical yield of iron (II) chloride if the reaction begins with 34
          grams of iron (II) bromide?
      c) What is the percent yield of iron (II) chloride if my actual yield is 4.0 grams?

5)   a) Balance the following equation: $C_3H_8 + O_2$ → $CO_2 + H_2O$
      b) If a reaction starts with 5.0 grams of C_3H_8, what is the theoretical yield of
          water?
      c) CHALLENGE: If the percent yield is 75%, how many grams of water were
          produced?