

## CHEM ELECTRONEGATIVITY PRACTICE

Using the electronegativity values in your notes...

- a) calculate the bond strength
- b) classify each bond as NONPOLAR COVALENT, MODERATELY POLAR COVALENT, VERY POLAR COVALENT, or IONIC.

Then

- c) name each compound
- d) classify each compound

- |   |   |
|---|---|
| 1) Ba <sub>3</sub> P <sub>2</sub>                   | 9) CCl <sub>4</sub>                                       |
| 2) Al <sub>2</sub> O <sub>3</sub>                   | 10) the Zn-C bond in Zn(CH <sub>3</sub> COO) <sub>2</sub> |
| 3) CaI <sub>2</sub>                                 | 11) NaF   |
| 4) H <sub>2</sub> O                                 | 12) O <sub>2</sub>  |
| 5) the Na-S bond in Na <sub>2</sub> SO <sub>4</sub> | 13) Cu <sub>3</sub> N                                     |
| 6) Br <sub>2</sub>                                  | 14) the K-Cl bond in KClO <sub>3</sub>                    |
| 7) N <sub>2</sub> O <sub>5</sub>                    | 15) SrS   |
| 8) RbBr   |   |