

## CHEMISTRY FORMULA PRACTICE #2

Write formulas for the following and *classify* as binary ionic (BI), binary molecular (BM), ternary ionic (TI), or other.

The highlighted formulas are covered in the binary molecular chapter.

- 1) potassium nitride
  - 2) potassium nitrate
  - 3) potassium nitrite
  - 4) aluminum hydroxide
  - 5) **diphosphorus trioxide**
  - 6) strontium cyanide
  - 7) ammonium fluoride
  - 8) lead(II) iodide
  - 9) tin(IV) sulfide
  - 10) sodium phosphate
  - 11) iron(II) chloride
  - 12) iron(III) perchlorate
  - 13) copper(I) hypochlorite
  - 14) zinc permanganate
  - 15) calcium phosphide
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Name the following and *classify* as binary ionic (BI), binary molecular (BM), and ternary ionic (TI), or other.

The highlighted formulas are covered in the binary molecular chapter.

- 16) NaF
- 17) Ca(NO<sub>2</sub>)<sub>2</sub>
- 18) K<sub>2</sub>Cr<sub>2</sub>O<sub>7</sub>
- 19) **SCl<sub>6</sub>**
- 20) Na<sub>3</sub>PO<sub>3</sub>
- 21) **As<sub>2</sub>Br<sub>5</sub>**
- 22) **NO<sub>2</sub>**
- 23) SnO
- 24) (NH<sub>4</sub>)<sub>3</sub>P
- 25) Al<sub>2</sub>(SO<sub>3</sub>)<sub>3</sub>
- 26) CaCrO<sub>4</sub>
- 27) Mg(C<sub>2</sub>H<sub>3</sub>O<sub>2</sub>)<sub>2</sub>
- 28) Cr(HCO<sub>3</sub>)<sub>3</sub>
- 29) **SO<sub>3</sub>**
- 30) Na<sub>2</sub>Se