

## CHEMICAL FORMULA PRACTICE #1

Name the following compounds and *classify* as binary ionic (BI), binary molecular (BM), ternary ionic (TI), or other (more than three symbols).

The highlighted formulas are covered in the binary molecular chapter.

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|--|----------------------------------|
| 1) $\text{CaF}_2$                              | 18) $\text{Al}(\text{ClO})_3$    |
| 2) $\text{Na}_3\text{N}$                       | 19) $\text{Mg}(\text{NO}_3)_2$   |
| 3) $\text{OF}_2$                               | 20) $\text{ZnCO}_3$              |
| 4) $\text{Rb}(\text{C}_2\text{H}_3\text{O}_2)$ | 21) $\text{PbS}$                 |
| 5) $\text{Mg}_3\text{P}_2$                     | 22) $\text{AgClO}_4$             |
| 6) $\text{N}_2\text{O}_5$                      | 23) $\text{SO}_2$                |
| 7) $\text{MnCrO}_4$                            | 24) $\text{Zn}(\text{CN})_2$     |
| 8) $\text{PCl}_5$                              | 25) $\text{N}_2\text{O}$         |
| 9) $\text{Cr}_3(\text{PO}_4)_2$                | 26) $\text{Sr}(\text{ClO}_2)_2$  |
| 10) $\text{Co}_2(\text{Cr}_2\text{O}_7)_3$     | 27) $\text{NH}_4\text{Br}$       |
| 11) $\text{SnO}$                               | 28) $\text{CdSO}_4$              |
| 12) $\text{NaHCO}_3$                           | 29) $\text{Mn}_2(\text{SO}_3)_3$ |
| 13) $\text{LiOH}$                              | 30) $\text{CaS}$                 |
| 14) $\text{CO}$                                | 31) $\text{AgNO}_3$              |
| 15) $\text{KClO}_3$                            | 32) $\text{Na}_2\text{SiO}_3$    |
| 16) $\text{SnO}_2$                             | 33) $\text{GaI}_3$               |
| 17) $\text{Cu}(\text{MnO}_4)_2$                | 34) $\text{OF}_6$                |

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Write the formula for the following compounds and *classify* as binary ionic (BI), binary molecular (BM), or ternary ionic (TI), or other.

The highlighted formulas are covered in the binary molecular chapter.

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|-------------------------------------|--------------------------------|
| 35) barium iodide                   | 43) calcium hydroxide          |
| 36) potassium hydroxide             | 44) iron(III) perchlorate      |
| 37) <b>diphosphorus pentaiodide</b> | 45) iron(II) nitrite           |
| 38) lead(II) nitrate                | 46) silver chloride            |
| 39) aluminum acetate                | 47) <b>carbon dioxide</b>      |
| 40) cesium hypochlorite             | 48) cobalt(II) chloride        |
| 41) ammonium sulfate                | 49) <b>dinitrogen monoxide</b> |
| 42) lithium nitride                 | 50) <b>sulfur hexafluoride</b> |