

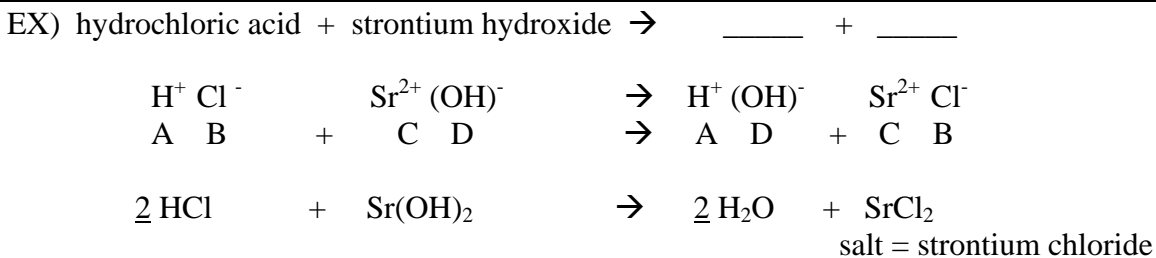
ACIDS, BASES, AND NEUTRALIZATION REACTIONS

Neutralization Reactions “ACID + BASE → WATER + SALT”

a) Write and balance these double displacement reactions. $AB + CD \rightarrow AD + CB$

- If the formula is not provided, you must “crisscross” to get it.
- Remember, to get the products, you must “un-crisscross” and “re-crisscross” the reactant ions. If you can do that mentally, you do not have to show your crisscross work.

b) All salts must be named. Use the complete polyatomic ion list if needed.



- 1) acetic acid + sodium hydroxide → _____ + _____
- 2) HNO₂ + calcium hydroxide → _____ + _____
- 3) H₃BO₃ + barium hydroxide → _____ + _____
- 4) HClO₂ + strontium hydroxide → _____ + _____
- 5) hydrochloric acid + sodium hydroxide → _____ + _____
- 6) HCN + aluminum hydroxide → _____ + _____
- 7) sulfuric acid + potassium hydroxide → _____ + _____
- 8) HBr + calcium hydroxide → _____ + _____
- 9) nitric acid + lithium hydroxide → _____ + _____
- 10) H₂C₂O₄ + barium hydroxide → _____ + _____
- 11) HClO₄ + sodium hydroxide → _____ + _____
- 12) H₂SO₃ + potassium hydroxide → _____ + _____
- 13) HF + strontium hydroxide → _____ + _____
- 14) HClO + aluminum hydroxide → _____ + _____
- 15) HMnO₄ + lithium hydroxide → _____ + _____
- 16) HClO₃ + barium hydroxide → _____ + _____
- 17) HI + potassium hydroxide → _____ + _____
- 18) phosphoric acid + calcium hydroxide → _____ + _____
- 19) H₂S + aluminum hydroxide → _____ + _____
- 20) H₃BO₃ + aluminum hydroxide → _____ + _____