

Bauck's CHEMISTRY Ch. 9 Test Review

This is an optional assignment due the day of the test.

Materials: loose leaf paper, pencil, calculator (You will be given a periodic table.)
Format: math problems: two-step mole problems, short gas density, long gas density, percent composition, empirical formula, long molecular formula
Test date: _____
Test value: 200 points

BACKGROUND INFO:

- 1) **Empirical formula**—What is it?
- 2) **GAM, GFM, GMM**—What are these? How are they measured?
- 3) What is the **molar volume** of a gas at STP?
- 4) **Molar mass**—What is it?
- 5) **Mole**—What is it? Why is it so important in chemistry?
- 6) **Molecular formula**—What is it?
- 7) **Percent composition**—What is it?
- 8) **Representative particles**—What are they? List the four main types of r.p.'s and give an example of each.
- 9) **STP**—What does this stand for? When is it used?

MATH PROBLEMS – Give an example of each for this review.

- 10) mol \rightarrow r.p. or r.p. \rightarrow mol
 - 11) mol \rightarrow g or g \rightarrow mol
 - 12) mol \rightarrow L or L \rightarrow mol (gas at STP)
 - 13) g \rightarrow r.p. or r.p. \rightarrow g
 - 14) g \rightarrow L or L \rightarrow g (gas at STP)
 - 15) r.p. \rightarrow L or L \rightarrow r.p. (gas at STP)
 - 16) mol \rightarrow number of atoms or ions in a compd. or number of atoms or ions in a compd. \rightarrow mol
 - 17) find molar mass (GFM or GMM)
 - 18) gas density: g/L \rightarrow g/mol g/mol \rightarrow g/L
 - 19) percent composition
 - 20) empirical formula
 - 21) molecular formula, short version
 - 22) molecular formula, long version
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*** Note ***

There will be at least one question pertaining to material in past chapter(s) or unit(s).