

Bauck's CHEMISTRY Ch. 13 Test Review

This is an optional assignment due the day of the test.

Materials: loose leaf paper, pen and/or pencil, calculator (You will be given a periodic table.)
Format: multiple choice, naming hydrates, molarity math problems, short answer essays
Test date: _____
Test value: 200 points

VOCABULARY

- 1) **Colligative properties** of solutions—Give examples.
- 2) **Colloid**—What is it; compare and contrast with **solution** and **suspension**
- 3) **Concentrated**—What is this; contrast to **dilute**.
- 4) **Dilute**—What is this; contrast to **concentrated**.
- 5) **Electrolytes**—What are they? List characteristics.
- 6) **Hydrates**—What are they? (Be able to name hydrates.) Give an example for this review.
- 7) **“Like Dissolves Like”**—Explain what this means and how it relates to forming solutions. Give an example.
- 8) **Molarity**—What does this measure? What is the molarity symbol and unit? (Be able to do molarity math problems. See next section below.)
- 9) **Nonpolar** substances—What are they? Give examples. Contrast with **polar** substances.
- 10) **Polar** substances—What are they? Give examples. Contrast with **nonpolar** substances.
- 11) **Saturated** solution—What is it? List characteristics. Contrast with **unsaturated** and **supersaturated** solutions.
- 12) **Solute**—What is it? Give examples.
- 13) **Solution**—What is it? Give examples. Compare and contrast with **suspension** and **colloid**.
- 14) **Solvent**—What is it? Give examples.
- 15) **Supersaturated** solution— What is it? List characteristics. Contrast with **unsaturated** and **saturated** solutions.
- 16) **Suspension**—What is it? Compare and contrast with **solution** and **colloid**.
- 17) **Unsaturated** solution— What is it? List characteristics. Contrast with **saturated** and **supersaturated** solutions.
- 18) **Water**—List three characteristics Why is it a good general solvent?

MATH PROBLEMS

$$M = \text{mol/L} \quad (\text{equation given on test})$$

For this review, give an example of each of the following. Show all work and units.

MOLARITY (M)

- 19) Solve for molarity... given moles solute and volume of solvent
- 20) Solve for molarity... given grams solute and volume solvent
- 21) Solve for mol solute... given M and volume solvent
- 22) Solve for g solute... given M and volume solvent