Bauck's CHEM Ch. 11 Test Review

This is a 10/10 optional assignment due the day of the test.

Materials: loose leaf paper, pen and/or pencil, calculator
(You will be given a periodic table.)
Format: math problems, writing/balancing equations
Test date: _____
Test value: 200 points

Practice math problems! This is a math-based test.

BACKGROUND INFO:

1) How do you know when to use **Avogadro's number (6.02 x 10^{23})** in calculations?
2) List the four types of **representative particles**. Give an example of each.
3) What is the **molar volume** of a gas at STP? (number of Liters?)
4) Why is "**molar mass"** the same as **grams per mole, GFM** and **GMM**?
5) MATH PROBLEMS (with example problems to look at from our notes)
   Include one example of each bulleted category for this review.
   - **Mol-Mol** (E1 & E2) mol A → mol B
   - **Mass-Mass** (E3 & E4) g of A → g of B
   - **Mass-Mole** or **Mole-Mass** (E5 & E6) g of A → mol of B mole of A → g of B
   - **Mass-Volume** or **Volume-Mass** (E7 & E8)
     g of A → L of B L of A → g of B
   - **Volume-Volume** (E9 & E10) L of A → L of B
   - **Mass-Particle** or **Particle-Mass** (E11 & E12)
     g of A → r.p. of B r.p. of A → g of B
   - **Limiting Reactant** (E13)
   - **Percent Yield** (E14 & E15)
   - **Particle-Volume** (worksheets) r.p. of A → L of B L of A → r.p. of B

6) There are chemical reactions on the test:
   There will be equations to balance.
   Some have to be constructed from words, without products given.
   Some will be given to you already balanced.

   For this review, give an example of a balanced equation, a skeleton (unbalanced) equation, and an equation in words which you construct and balance.

*** Note ***
There will be at least one question pertaining to material in past chapter(s) or unit(s).